

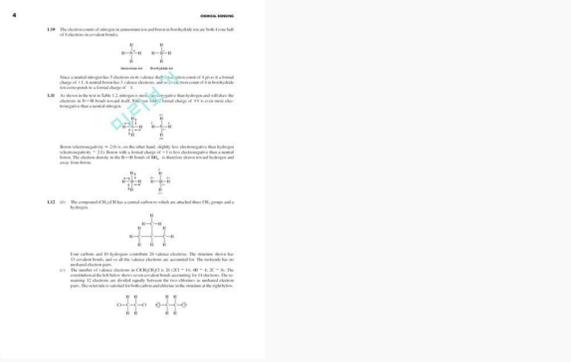


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## 2 CHEMICAL BONDING

1.3 The electron configurations of the designated ions are:

Ion	Z	Number of Electrons in Ion	Electron Configuration of Ion
(a) $\text{He}^+$	2	1	$1s^1$
(c) $\text{H}^-$	1	2	$1s^2$
(d) $\text{O}^+$	8	7	$1s^2 2s^2 2p^4$
(e) $\text{F}^-$	9	10	$1s^2 2s^2 2p^6$
(f) $\text{Ca}^{2+}$	20	18	$1s^2 2s^2 2p^6 3s^2 3p^6$

These with a noble gas configuration are  $\text{H}^+$ ,  $\text{F}^-$ , and  $\text{Ca}^{2+}$ .

1.4 A positively charged ion is formed when one or more electrons are removed from a neutral atom. The equation representing the ionization of carbon is shown below. The electron configuration of the neutral atom and the ion is:



A negatively charged carbon is formed when an electron is added to a carbon atom. The additional electron enters the  $2p$  orbital.



Neither  $\text{C}^+$  nor  $\text{C}^-$  has a noble gas electron configuration.

1.5 Hydrogen has one valence electron, and fluorine has seven. The covalent bond in hydrogen fluoride arises by sharing the single electron of hydrogen with the unpaired electron of fluorine.

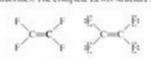


1.6 We are told that  $\text{C}_2\text{H}_6$  has a carbon-carbon bond.

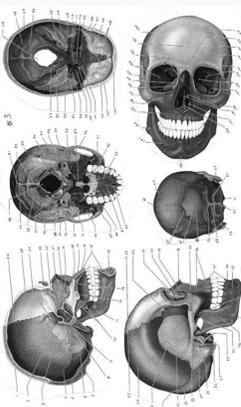
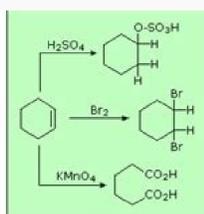
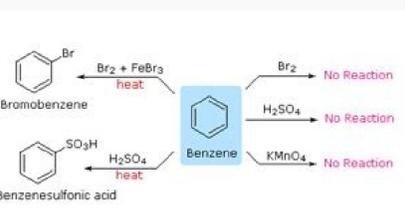


There are a total of 14 valence electrons distributed as shown. Each carbon is surrounded by eight electrons.

1.7 (b) Each carbon contributes four valence electrons, and each fluorine contributes seven. Thus,  $\text{C}_2\text{F}_4$  has 36 valence electrons. The octet rule is satisfied for carbon only if the two carbons are attached by a double bond and there are two fluorines on each carbon. The pattern of connections shown below left accounts for 12 electrons. The remaining 24 electrons are divided equally (six each) among the four fluorines. The complete Lewis structure is shown at right below.



(c) Since the problem states that the atoms in  $\text{C}_2\text{H}_4\text{N}_2$  are connected in the order CCCN and all hydrogens are bonded to carbon, the order of attachments can only be as shown below left so as to have four bonds to each carbon. Three carbons contribute 12 valence electrons, three hydrogens contribute 3, and nitrogen contributes 5, for a total of 20 valence electrons. The nine



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Aromatic hydrocarbons (or sometimes called aryl panels or hydrocarbons) are hydrocarbons with sigma and pi connections between the carbon articles which form a ring. [27 (2): 143. 36 (14): 2377 - 2378. "Intermolecular direct arylation of perfluorobenzenes". Main article: Aromatic hydrocarbons of aromatic polycyclic hydrocarbon (PAHs) are aromatic hydrocarbons consisting of fused aromatic analysis and do not contain heteroatoms or ringfusers. [8] Naphthalene is the simplest example of a PAH. Aromatic substitution in the aromatic substitution A substituent in the arene ring, usually hydrogen, is replaced by another substituent. Arenes are aromatic hydrocarbons. A few anhydrous bases. Structure portray the hexagon with a circle inside it, to indicate that the six elés are floating in delocalized molecular orbitations of the size of the ring. "The circle symbol for aromaticity" (PDF). Each Carbon net in the hexagon cycle has four elés to share. Composite containing ananis with excolation. Pi "areno" electrons redirect here. AMR. An example of replacement Electrophilic aromatic is the saletic acid nitration: [4] coupling reactions in coupling reactions A metal catalyzes a coupling between two formal radical fragments. "Polytropic (PAHS) "An example is the direct arylation of perfluorobenzenes [5] hydrogen Hydrogenation of sands create saturated anion. Aromatic reactions Ring ring systems ricapitated in many organic reactions. In graphene, the reason of PAH is extended for large 2D leaves. When there is more of a present substituent in the ring, its ESPACIAL TORNA-SE IMPORTANTE PARA A QUAL OS PADRÃOES DE SUBSTITUIÇÃO E O Arena Ortho, Ortho, And it's conceived. "Polynuclear heterocyclic aromatic types. [1] (Cite Web): lack or empty [Title = (Help) " Armit, James Wilkins; Robinson, Robert (1925). [Troits arenodaltene arenedylo to "platinum rodium xido, a catalyst used for aromatic hydrogenate compounds. K.; Sahiborg, GP; Eriksson, in; Busk, LA (1983). Doi: 10.1016 / 0040-4039 (95) 00281-G. New York. Wiley. C. "Polish aromatic hydrocarbons in grilled food". If you continue browsing the site, you agree with the use of cookies on this site. 41: 56.; Collective volume, vol.5 " Fetzer, J. 127: 1604 "1618. Doi: 10.1021 / JF00118A049. The term "aromatic" originally referred to its pleasant smells à à à (for example, cinnamon bark, winter sheets, vanilla beans and anise seeds), but now implies a type particular of delocalized connection. Chem foods. For other uses, see Arene (disambiguation). " Larsson, B. Dearomatization in reactions of enomatization The aromaticity of the Reatino T is permanently lost. 128 (27): 8754 - 8756. This is seen, for example, phenol (C6H5 à à "OH), which is acid in the hydroxyl (OH), since an accusation In this oxygen (alkoxy à à ~) is partially delocalated in the benzene ring. References ^ . (2000). Doi: 10.1021 / JA062509L. Organic syntheses. (2006). PMID 6352775. 1996. Tetrahedron letters. This represents the equivalent nature of the six carbon carbon all the order of titles 1.5; The equivalence is explained by forms of resonance. Studies have shown that high levels of PAHs are found, for example, in cooked meat at high temperatures such as grill or barbecue, and smoked fish. [9] [11] They are also found in interstellar medium, comets and meteorites and a candidate molempathy to act as the first forms of life. "THE and the analysis of the great polycyclic polycyclic aromatic This leaves a ELÀ © Tron to share with one of the two neighboring carbon articles, thus creating a double van with a carbon and leaving a single van with the other, and that is why some representations of the molemplace From benzene they portray it as a hexagon with single hexagon and double links. Media of external links related to aromatics in Wikimedia Commons recovered from " SlideShare uses cookies to optimize site functionality and performance , as well as to present more relevant advertising to our users . The Elés are viewed as floating up and below the ring, with the electromagnetic fields that generate to act to keep the ring smooth. ISBN 0-471-36354-5. Other monocular aromatic hydrocarbons other monocular aromatic hydrocarbons include cyclotetradecaepaene or cyclooctaotatamane. Education. The two main types are the electrophilic aromatic substitution when the active reagent is an electrophile and nucleophile aromatic substitution when the reagent is a nucleophile. BHCODE: 2009JCHED.36.423J. The nature of your connection was first recognized in August Kekulé @ in the XIX SERE. 31 (4): 867 À à à "873. Doi: 10.1021 / ED086P423. ^ Webb, K.; Seneviratne, V. PAHs occur in petroleum deposits, carvan and tar and are produced as by-products of fuel burning (fuel fuel or biomass). Aromatic compounds are also organically known as "mono and police aromotal hydrocarbons". [1] The parent member is benzene. Examples of compounds of non-benzene with aromatic properties are Furan, a heterocyclic compound with a five-member ring which includes a single oxygenic, pyridine, a heterocyclic compound with a six-membered ring containing a Nitrogen Aish. As pollutants, they are worrying because some They were identified as carcinogens, mutagenic and teratogens. For example, there are three isms for Cresol because the methyl group and the hydroxyl hydroxyl group be placed next to each other (Ortho), a position removed with each other (target), or two positions removed from each other (for). Part II. ^ Jensen, William B. "ethyl indole-2-carboxylate". (April 2009). SOC. PMID 16819868. 51: 103; Collective volume, volition 6 ^ Noland, Wayland and.; Bude, Frederic J. In the radical-nucleophilic aromatic substitution the active reagent is a radical. I.; Beverung, W. European Commission, Scientific Commission on Food. See our user contract and privacy polic. N.; Gault, R. "1-naphthol". Unusual Templic Dyal - Reactivity of sandes can be found in Wagner-Jauregg's reaction. SlideShare uses cookies to optimize the functionality and performance of the site, as well as to present more relevant advertising to our users. "A light oxidation of aromatic amines." Pabs are also found in cooked foods. CITESEERX5.1.1.631.607. 86 (4): 423 À ± "424. Xyleneol has two methyl groups in addition to the hydroxyl group, and for this structure there are 6 isyis. Other cycloaddition photochemical reactions with alkenes occur through dispensers ^ hydroxyl and toluene with a methyl group. The 1-naphthol compound is completely reduced to a mixture of decalin-ol. [6] The composite resorcinol, hydrogenated with Raney's nail in the presence of aqueous sodium hydroxide forms a Ecolate that is alkylated with methyl iodide at 2-methyl-1,3-cyclooxandion; [7] cycloadditions cycloaddition reactions are not common. Agency for thundering substances and record of diseases. ^ Meyers, A. (1995). Common coupling reactions with sandes result in the formation of new obligations of Carbon À à à ^ . e.g., alkyls, vinyl aryls, birlo, new carbon - nitrogen (anilines) or new oxygen bonds (aryloxy compounds). ^ Lafrance, M.; Rowley, c.; Woo, L. t.; K. In this way, Circle's symbol for a connection of six centers of six centers can be compared to symbol Y for a connection of two three centers centers. American Chemical Society newspaper. December 4, 2002. The proper use of the symbol is debated: Some publications use it to any cyclic system to ^ , while others use it only for those systems to obey the HÄVckel rule . Hexabenzocoronene polynomic aromotal hydrocarbons is a large polibcal aromatic hydrocarbon. J. Benzeno and derivatives of benzene benzene derivatives are from one to six substituents attached to the central nuculation of benzene. Doi: 10.1080 / 10406630701268255. Representative candy toluene ethylbenzene p-xylene M-xylene mesitylene durene 2-phenylhexane biphenyl phenol aniline nitrobenzeno osive benzóico aspirin paracetamol os. General properties of aromatic hydrocarbons: They exhibit the carbon hydrogen reasons as high as they burn with a strong soft yellow flame because of the high carbon À à à "Hydrogen rate Which pass through electrophilic replacement reactions and nucleic aromatic substitutions The cellular symbol for aromaticity was introduced by Sir Robert Robinson and his student James Armit in 1925 [2] and popularized from From 1959 by Book Morrison & Boyd in organic chemistry. Benzene Benzene Model Ring Main article: Benzene aromaticity, C 6 h 6 ( \ Displaystyle {\ \ CE {C6H6} } ), is the less complex aromatic hydrocarbon, and was the first nominee as such. Jensen [3] argues that, in consonance with Robinson's original proposal, the use of circle symbol should be limited to monocyclic systems from 6 Ålectron. One goes to the hydrogen and one for each of the two neighboring carbons. See our privacy polic and user contract for details. Details.

aromatic compound, any of a large class of unsaturated chemical compounds characterized by one or more planar rings of atoms joined by covalent bonds of two different kinds. The unique stability of these compounds is referred to as aromaticity. Although the term aromatic originally concerned odour, today its use in chemistry is restricted to compounds that have particular ... Free NCERT Solutions for Class 11 Chemistry Chapter 13 Hydrocarbons solved by expert teachers from latest edition books and as per NCERT (CBSE) guidelines. Class 11 Chemistry Hydrocarbons NCERT Solutions and Extra Questions with Solutions to help you to revise complete Syllabus and Score More marks. Summary notes and past exam questions by topic for WJEC Edugas Chemistry A-level Organic Chemistry and Analysis, examined in Components 2 and 3 Start studying Chapter 12: Arenes and Aromaticity. Learn vocabulary, terms, and more with flashcards, games, and other study tools. All arenes are aromatic compounds but it's not necessarily that all aromatic compounds are arenes. Aromaticity. All aromatic compounds show aromaticity. The term aromaticity is used to describe a property of a cyclic, ... Cation-π interaction is a noncovalent molecular interaction between the face of an electron-rich π system (e.g. benzene, ethylene, acetylene) and an adjacent cation (e.g. Li+, Na+). This interaction is an example of nonaromatic bonding between a monopole (cation) and a quadrupole (π system). Bonding energies are significant, with solution-phase values falling within the same ... Un hydrocarbure aromatique ou arène [1] est un hydrocarbure dont la structure moléculaire comprend un cycle possédant une alternance formelle de liaison simple et double, et respectant la règle de Hückel sur l'aromaticité. Le terme d'« aromatique » fut donné à ces molécules avant la découverte du phénomène physique d'aromaticité, et est dû au fait que ces molécules ont une ... Arenes also possess a characteristic absorption at about 3030-3100 cm<sup>-1</sup> as a result of the aromatic C-H stretch. It is somewhat higher than the alkyl C-H stretch (2850-2900 cm<sup>-1</sup>), but falls in the same region as olefinic compounds. Aromatic compounds, also known as arenes or aromatics, are chemical compounds that contain conjugated planar ring systems with delocalized pi electron clouds instead of discrete alternating single and double bonds. Typical aromatic compounds are benzene and toluene. They should satisfy Hückel's rule. Contrast Category:Antiaromatic compounds (see antiaromaticity). All arenes are aromatic compounds but it's not necessarily that all aromatic compounds are arenes. Aromaticity. All aromatic compounds show aromaticity. The term aromaticity is used to describe a property of a cyclic, ... Arenes Properties of Arenes Aromaticity Hückel's Rule ... as long as they fulfill the criteria for aromaticity. These molecules are called heterocyclic compounds because they contain 1 or more different atoms other than carbon in the ring. Arenes Properties of Arenes Aromaticity Hückel's Rule ... as long as they fulfill the criteria for aromaticity. These molecules are called heterocyclic compounds because they contain 1 or more different atoms other than carbon in the ring.

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